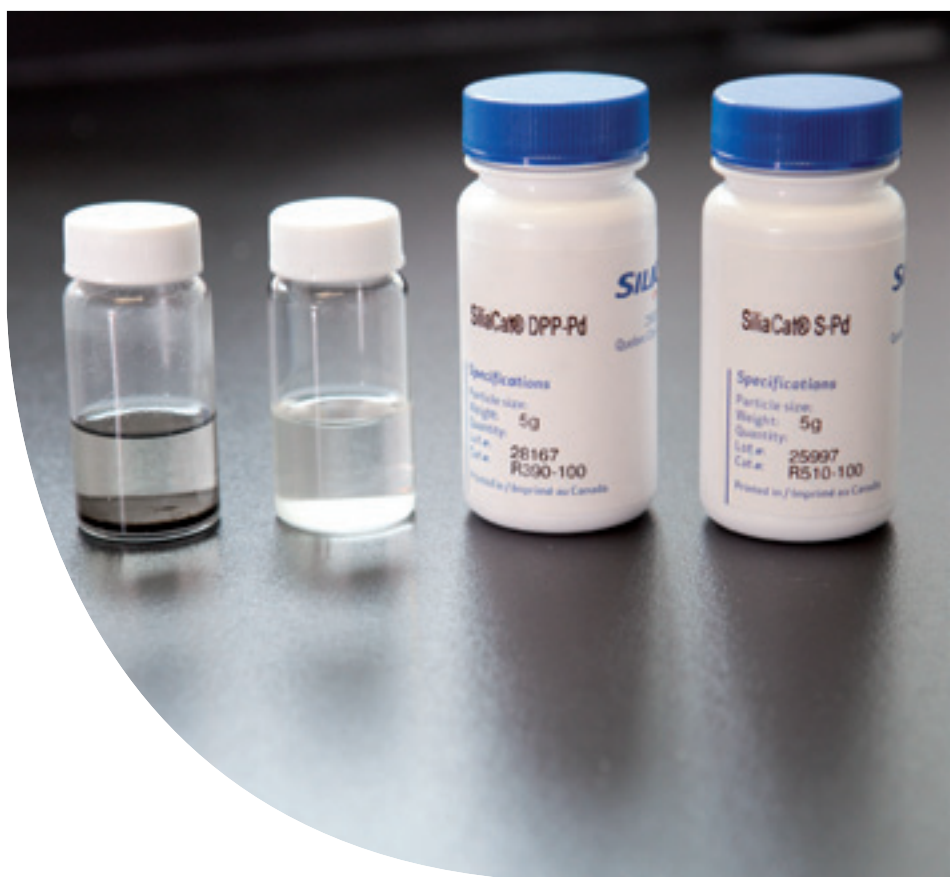


# SiliaCat<sup>®</sup>

## Heterogeneous Catalysts



Distributed by

**Greyhound Chromatography and Allied Chemicals**  
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## Catalytic Reactions with SiliaCat®

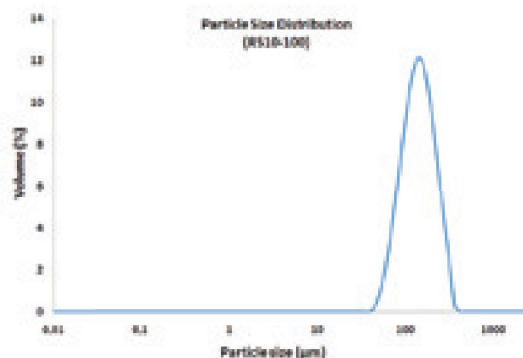
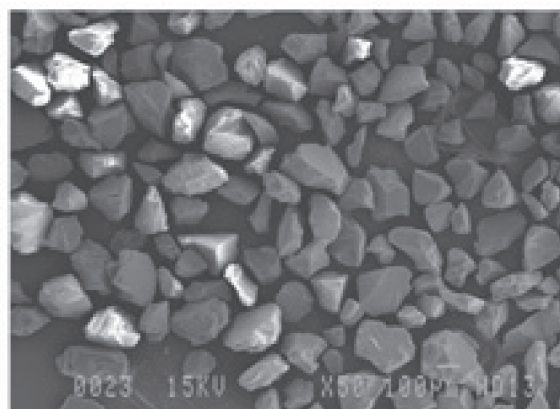
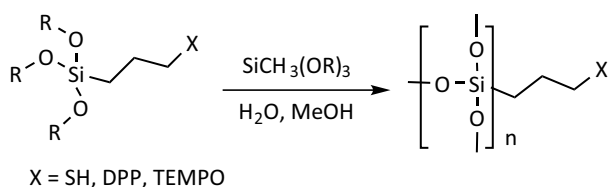
Advantages of using SiliaCat® heterogeneous catalysts over competitive products include:

- High stability
- Rigid & Porous Structure (*no swelling*)
- Compatibility with a wide range of solvents
- Ease of use: no swelling or static charge
- Leach-proof
- High turnover number (*TON*)
- Fast kinetics
- Accurate loading



## The SiliaCat Matrix

Inspired by the ORganically MOdified SILica (*ORMOSIL*) technology, the SiliaCat family is composed by new and innovative catalysts. Resulting from the co-condensation of two organosilane monomers by the sol-gel process (*confer condensation mechanism below*), the hybrid organic-inorganic materials present the highest stability and reactivity available with heterogeneous catalysts. Furthermore, the high cross linked framework presents an unmatched resistance, significantly better than the usual post-synthesis functionalized ligand.



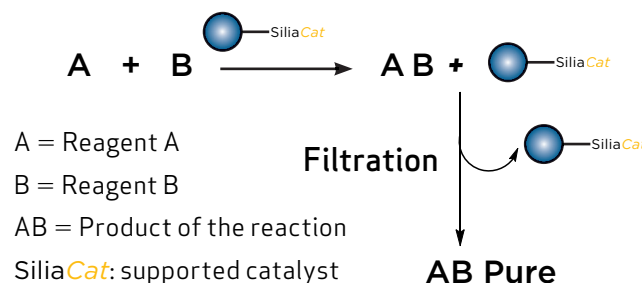


## What are SiliaCat Heterogeneous Catalysts?

Usually, heterogeneous catalysts supported on a silica matrix are immobilized by post-modification of the inorganic support. These supports, however, present a high degree of leaching due to the poor stability of the immobilized phase. With SiliaCat Heterogeneous Catalysts, the ligand is directly cross linked in an organic-inorganic framework. This results a high degree of stability of the catalysts. Compared to homogeneous catalysts, SiliaCat exhibits a good reactivity and selectivity with one of the major advantages being that the catalyst is eliminated from the reaction mixture by a simple filtration. Forget your purification problems with our SiliaCat catalysts family.

The process for using SiliaCat Heterogeneous Catalysts is outlined in the scheme below.

What is SiliaCat Heterogeneous Catalyst?





## Features and Benefits of SiliaCat Catalysts

Features & Benefits of SiliaCat	
Features	Benefits
Inertness within entrapped molecules	High conversion and yield
Reagent concentrated at the surface of the material	Reliable and reproducible synthesis
Robustness	High thermal and mechanical stabilities
Rigid and porous structure	No swelling, solvent independent and air stable Conditions do not have to be inert
Leach-proof organoceramic matrix	No contamination of APIs
High and accurate catalyst loading	Less catalyst required over competitive products
High turnover number (TON)	Catalytic amount (< 1 mol %)
Reusability	Multi-uses possible
Ease of handling and purification	Free flowing, no static charge Easily removed by simple filtration
Ease of scalability	Scalable from mg up to multi-ton scale
Flexible formats	Amenable to use in SiliaSep & SiliaPrep Cartridges
Available in bulk quantities	Can be delivered in large quantities and always in stock

## SiliaCat Heterogeneous Catalysts Product Range

SiliCycle, a leader in functionalized silica gels, has developed various catalysts at competitive prices.

SiliaCat Heterogeneous Catalysts Portfolio*			
SiliaCat Name	Product Number	Structure	Brief Description
SiliaCat DPP-Pd	R390-100		The significant costs associated with precious metal catalysts and their tendency to remain in organic products has generated interest for solutions that increase reactivity and can enable the recovery and reuse of these metals. SiliaCat DPP-Pd is a unique diphenylphosphine palladium (II) heterogeneous catalyst made from a leach-resistant organoceramic matrix.
SiliaCat S-Pd	R510-100		The significant costs associated with precious metal catalysts and their tendency to remain in organic products has generated interest for solutions that increase reactivity and can enable the recovery and reuse of these metals. SiliaCat S-Pd is a unique thiol-based palladium (II) heterogeneous catalysts made from a leach-resistant organoceramic matrix.
 SiliaCat Pd <sup>0</sup>	R815-100		SiliaCat Pd <sup>0</sup> is a new series of patent-protected sol-gel-entrapped Pd nanocatalysts. It is made from highly dispersed Pd nanoparticles ( <i>uniformly in the range 4.0–6.0 nm</i> ) encapsulated within an organosilica matrix.
 SiliaCat Pt <sup>0</sup>	R820-100		SiliaCat Pt <sup>0</sup> is made of organosilica physically doped with nanostructured platinum (0), and is both stable and efficient. This catalyst was successfully prepared by a novel and simple sol-gel route. In the new procedure, Pt nanoparticles ( <i>uniformly in the range 1.7–3.15 nm</i> ) are encapsulated via an alcohol-free sol-gel process typical of enzyme sol-gel encapsulation.
SiliaCat TEMPO	R723-100		SiliaCat TEMPO is a new oxidizing catalyst made from a proprietary class of organosilica-entrapped radicals. This encapsulation process confers enhanced reactivity and properties. The leach-resistant organoceramic matrix makes SiliaCat TEMPO highly efficient and selective compared to homogeneous TEMPO reagents. It also has a superior performance compared to polymer-supported TEMPO and silica-supported TEMPO in terms of both selectivity and stability. With SiliaCat TEMPO, no activation is required prior to use and selective aldehyde vs acid oxidation is possible. <i>U.S. Patent: 6,797,773 B1,2004</i>

Formats: 5g, 10g, 25g, 50g, 100g, 250g, 500g, 1kg, 5kg, 10kg, 25kg, ...



## SiliaCat Heterogeneous Catalysts Portfolio

Typical Applications	SiliaCat Typical Characteristics						SiliaCat Name
	Color	Endcapping	Molecular Loading	Typical Tap Density	Solvent Compatibility	Prolonged Storage	
Suzuki, Heck Sonogashira, Kumada, Stille	Yellow	Yes	≥ 0.2 mmol/g	415 g/L	All solvents, aqueous and organic	Keep dry	SiliaCat DPP-Pd
Suzuki, Heck Sonogashira, Kumada, Stille	Red - Orange	Yes	≥ 0.3 mmol/g	550 g/L	All organic solvents	Keep dry	SiliaCat S-Pd
Suzuki, Heck Sonogashira, Kumada, Stille, Selective debenzylation, Selective hydrogenation	Dark brown to black	Yes	-	-	All solvents, aqueous and organic	Keep cool (< 8 °C) Under Argon	SiliaCat Pd <sup>0</sup>
Selective reduction of nitroarenes, Hydrosilylation	Dark brown to black	Yes	-	-	All solvents, aqueous and organic	Keep dry Under Argon	SiliaCat Pt <sup>0</sup>
Oxidation of alcohols or Aldehydes	Orange	Yes	≥ 0.70 mmol/g	639 g/L	All solvents, aqueous and organic	Keep dry	SiliaCat TEMPO


 CONFIDENTIALITY  
ASSURED

## Catalyst Screening Service

Looking for the right SiliaCat Heterogeneous Catalyst to use? Contact us to take advantage of SiliCycle's expertise in catalysis. Our R&D team can find the optimal conditions for you.

Our Catalyst Screening Service is innovative because it provides a turn key solution to the pharmaceutical and manufacturing industries. Working with the substrates you identify, SiliCycle's chemists will quickly develop the most efficient catalysis process (*which catalyst and solvent to use, optimal concentrations, etc.*).

We guarantee confidentiality, since in most cases our screening service requires us to work with APIs or other patented materials. This will ensure an easy technology transfer.

Make the call many major pharmaceutical companies have already made. Contact us to discuss how we can help you to reach your goals. Many screening services are available to fit your needs and budget.

## SiliaCat - Regulatory Information

SiliaCat Heterogeneous Catalysts are being used more and more by GMP pharmaceutical, biotechnology, and fine chemical industries as well as contract research and manufacturing organizations. Many have run their own analysis proving SiliaCat can safely be used without compromising the purity of their compounds due to leaching.

Need specific regulatory files? SiliCycle can work with you to fulfill your requirements. We can provide custom regulatory documentations that include specific analytical tests in line with your needs.

SiliCycle is committed to high quality standards and strives to provide default-free products. In doing so, all products are manufactured in an ISO 9001:2008 compliant facility and submitted to a stringent quality control. Every lot must meet the quality specifications to be released for sale and a sample from every batch is kept for subsequent analysis. All products are shipped with the following information:

- Certificate of Analysis
  - Purity (*Leachables and extractables*)
  - Molecular loading
  - Surface Coverage
  - Volatile Content
- Material Safety Data Sheets (*MSDS*)
- BSE/TSE Declaration (*no animal origin*)
- Relevant Technical Information



## Experimental Procedures and Optimization

Typical experimental procedures can be found for each catalytic reaction. Please note that these procedures are the starting suggestions meant to be starting points. Sometimes, optimization steps need to be

undertaken to optimize yields and increase selectivity. Various parameters can be changed, one at a time or simultaneously, to improve results.

### Number of mol % of SiliaCat Catalysts

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For each new experiment, we suggest using a molar percent of SiliaCat with respect to the substrate. This quantity has to then be optimized in order to obtain a good catalytic activity with the lowest consumption of the SiliaCat. For initial experiments we suggest to use an higher mol % of SiliaCat Catalyst in respect to the substrate and then decrease the quantity if yield and kinetics are already in line with your needs. During development applications work at SiliCycle, we always start using 1 mol % of catalyst.

### Solvent

---

SiliaCat can safely be used in a wide range of organic and aqueous solvents commonly used in laboratory and in process work, such as DMF, DMSO, THF, ACN, alcohols, ethers, chlorinated solvents, water, etc. The nature of the solvent does sometimes influence the catalytic efficiency, however. If yield is low or kinetics are too slow, changing solvent or adding a co-solvent should be considered.

### Solution Concentration

---

At low substrate concentration, the activity of the catalyst will be directly proportional to the number of moles of substrate available. If you increase the concentration of the substrate, the activity will increase until the active sites become saturated. So the substrate concentration is a parameter that needs to be optimized to develop your catalytic conditions.

### Temperature

---

A catalyst's purpose is to enhance the kinetics of a reaction, so we recommend running the experiments at room temperature. In the optimization step, the temperature could be adjusted, if it is needed.

### Reaction Time

---

In the case that the TOF is low, and increasing the temperature to increase the kinetics is not possible, we recommend increasing the contact time with the catalyst to complete the reaction. Also, in this case, increasing of the amount of catalyst is an option.

## SiliaCat's Compatibility with New Technologies

### SiliaCat In Flow Chemistry and Microwave Assisted Experiments

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SiliaCat can also be used in flow chemistry and under microwave radiation. In flow chemistry, simply place the SiliaCat inside the solid-phase reactor included in the flow system (*i.e.* Syrris Asia® Solid Phase Chemistry Reactor) and run the reaction. See page 78 for more details.

In microwave experiments, SiliaCat showed excellent catalytic efficiency in a short period of time. See following pages for the different applications developed.

## Catalysis Definitions and Calculation

SiliaCat Heterogeneous Catalysts are sol-gel silica-supported catalysts that can be used to replace homogeneous catalysts. The process for using SiliaCat is outlined in the scheme page 21.

### What is a Turnover Number (TON)?

---

In catalysis, the term turnover number has two meanings: the number of moles of substrate that a mole of catalyst can convert before becoming inactivated and is the amount of substrate converted per the amount of catalyst used.

In theory, the Ideal catalyst would have an infinite turnover number and would never be consumed. In practice, turnover numbers begin at 100 and can go up to a million, more so in some cases.

### What is a Turnover Frequency (TOF)?

---

A catalyst's turnover frequency number, or turnover number per time unit, characterizes its level of activity. So the TOF is the total number of moles transformed into the desired product by one mole

of active site per hour. The larger the TOF, the more active the catalyst.

$$\text{TOF} = \text{TON}/\text{hour}$$

### How to Calculate the Amount of SiliaCat Needed Based on Mol %?

---

One mol % (1 mol %) means 0.01 molar equivalent. If 3 mmol of the substrate is used, then 0.03 mmol of SiliaCat catalyst is required. To determine the weight

of the catalyst needed, simply divide this value by the loading of the catalyst. For example, SiliaCat DPP-Pd typical loading is 0.2 mmol/g, so 0.15 g is needed.

### Solvent Molar Concentration

---

A 1.2 M solvent concentration means:

1.2 mmol of substrate per mL of solvent  
(or 1.2 mol of substrate per L of solvent)

$$\text{Volume of solvent needed} = \frac{\text{mmol of substrate used}}{\text{molar concentration desired}}$$

For example, if 3 mmol of the substrate is used, then, 2.5 mL of solvent is necessary to reach a 1.2 M concentration.

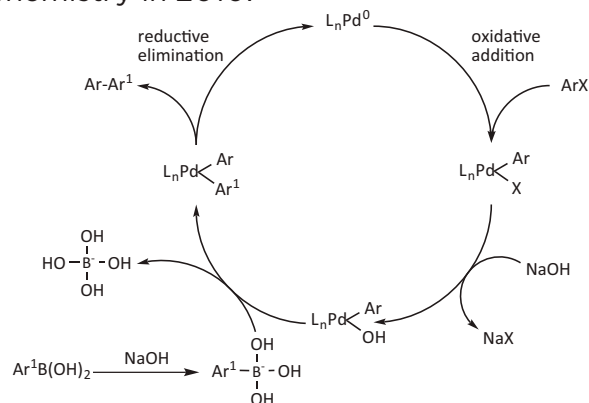




## Suzuki Coupling Using Pd-based SiliaCat

The Suzuki coupling (*also called Suzuki-Miyaura reaction*) is the reaction between a boronic acid and a halide catalyzed by a palladium (0) catalyst widely used in organic synthesis. At first, only aryl and vinyl substrates could undergo Suzuki coupling. Now, catalysts are becoming so powerful that the substrate scope has broadened to include: alkyl-, alkenyl- & alkynyl- halides, triflates and organoboranes, trifluoroborates or borate esters.

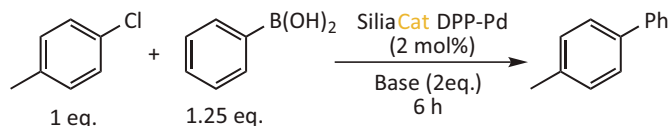
Its discovery was awarded the Nobel Prize in Chemistry in 2010.



### Solvent and Base Effects

The choice of solvent and base play an important part in the Suzuki reaction. Different solvents and bases were tested to find the most suitable combination. Total conversion was obtained in both ethanol and propanol. With THF, dioxane, toluene and DMF, the kinetics were lower.

For the base, potassium carbonate ( $K_2CO_3$ ) is the best. However, in some cases,  $Na_2CO_3$  and  $NaOAc$  can also be used.



#### Solvent and Base Effects

Solvent	Temp. (°C)	Conversion / Selectivity (%)					
		$K_2CO_3$	$Na_2CO_3$	KOAc	NaOAc	$K_2HPO_4$	$Et_3N$
MeOH	64	74 / 95	69 / 99	63 / 98	63 / 98	73 / 100	72 / 93
EtOH	77	100 / 98	100 / 97	82 / 99	85 / 100	79 / 100	77 / 93
EtOH/H <sub>2</sub> O (15%)	77	100 / 100	82 / 100	78 / 100	88 / 100	86 / 98	89 / 95
1-PrOH	90	100 / 95	70 / 97	90 / 99	91 / 99	15 / 100	20 / 95
2-PrOH	77	100 / 100	43 / 93	90 / 99	72 / 100	50 / 100	20 / 100
THF	64	30 / 93	15 / -	45 / 89	35 / 94	37 / 95	5 / -
MeTHF	77	40 / 95	33 / 100	39 / 100	56 / 100	30 / 97	4 / -
Dioxane	90	50 / 90	30 / 93	56 / 93	35 / 94	20 / 90	No reaction
Toluene	90	47 / 98	23 / 87	49 / 96	10 / 90	65 / 95	No reaction
DMF	90	50 / 100	30 / 100	15 / 100	17 / 100	7 / 100	No reaction

### Catalyst Concentration Effect

Decreasing the mol % of the catalyst lowers the kinetics of the reaction, but the total conversion can still be achieved. In this example, the addition of water significantly improves catalyst activity. Even if the catalyst amount is divided by 10, the TOF is still increased by a factor of five.

(Cdns: SiliaCat DPP-Pd,  $PhB(OH)_2$  (1.1 eq.),  $K_2CO_3$  (1.5 eq.) RT).

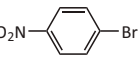
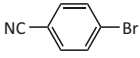
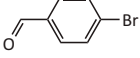
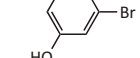
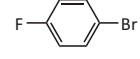
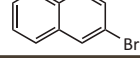
#### Catalyst Concentration Effect

mol %	Solvent (M)	Time (h)	Conv. (%)	TON	TOF
0.2	EtOH (0.05)	0.5	100	500	1,000
0.1	EtOH (0.05)	1	100	1,000	1,000
0.01	EtOH/H <sub>2</sub> O (0.08)	2	100	10,000	5,000
0.002	EtOH/H <sub>2</sub> O (0.08)	16	100	50,000	3,125

## Pd-Based SiliaCat's Catalytic Performance Comparison and Reusability

All SiliaCat Pd-based catalysts can be used for Suzuki coupling. The table below presents the best conditions for bromo- substrates. It can be seen that even with half the catalyst amount, SiliaCat Pd<sup>0</sup> is the more active catalyst.

For substrates with electron-withdrawing groups, SiliaCat catalysts can be reused more than 5 times with a minimal loss of activity and leaching. For substrates containing an electron-donating group, SiliaCat catalysts can be used up to 3 times with only a small effect on activity.

Pd-Based SiliaCat's Catalytic Performance Comparison and Reusability								
Substrate (R)	SiliaCat Performance Comparison [Conversion / Selectivity (%)]			Reusability [Conversion / Selectivity (%)] Pd & Si Leaching (ppm) <sup>1</sup>				
	DPP-Pd (1 mol %) <sup>a-b</sup>	S-Pd (1 mol %) <sup>c</sup>	Pd <sup>0</sup> (0.5 mol %) <sup>d</sup>	Run 2	Run 3	Run 4	Run 5	
Electron-Withdrawing		100 / 100 Pd: 0.1, Si: 2	100 / 99	100 / 99	100 / 100 Pd: 0.05, Si: 1	100 / 100 Pd: 0.08, Si: 1.5	100 / 100 Pd: 0.1, Si: 3	99 / 98 Pd: 0.1, Si: 3.5
		100 / 97 Pd: 0.1, Si: 3	100 / 99	99 / 97	98 / 99 Pd: 0.1, Si: 8	98 / 99 Pd: 0.07, Si: 5	100 / 99 Pd: 0.1, Si: 6	99 / 98 Pd: 0.1, Si: 5
		100 / 97 Pd: 0.1, Si: 6	94 / 88	95 / 98	99 / 90 Pd: 0.2, Si: 7	97 / 92 Pd: 0.2, Si: 8	99 / 98 Pd: 0.1, Si: 4	98 / 97 Pd: 0.1, Si: 5
Electron-Donating		100 / 99 Pd: 0.9, Si: 5	82 / 100	83 / 100	100 / 100 Pd: 0.6, Si: 9	100 / 98 Pd: 0.4, Si: 7	60 / 97 Pd: 0.05, Si: 6	-
		100 / 80 Pd: 0.07, Si: 3	94 / 100	98 / 99	99 / 99 Pd: 0.04, Si: 1.5	98 / 98 Pd: 0.1, Si: 2	81 / 94 Pd: 0.06, Si: 2	73 / 95 Pd: 0.03, Si: 7
		100 / 99 Pd: 2.1, Si: 10	72 / 95	97 / 95	88 / 90 Pd: 0.3, Si: 7	75 / 95 Pd: 4, Si: 9	87 / 99 Pd: 0.6, Si: 10	68 / 96 Pd: 0.4, Si: 11

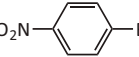
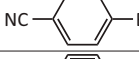


<sup>a</sup> Corresponds to "Run 1" in the reusability study.

General exp. cond.: 1 eq. substrate, 1.2 eq. PhB(OH)<sub>2</sub>, 2 eq. K<sub>2</sub>CO<sub>3</sub>; <sup>b</sup> MeOH (0.1 M), 2 h, 65°C; <sup>c</sup> EtOH/H<sub>2</sub>O (0.12 M) 4h, 77°C; <sup>d</sup> EtOH (0.12 M) 2h, 77°C.

<sup>1</sup> Using SiliaCat DPP-Pd as catalyst under the same conditions previously described. Run #1 is the result presented in the performance comparison section of the table.

The performance of the SiliaCat DPP-Pd and S-Pd catalysts for Suzuki coupling was also compared in microwave assisted experiments for brominated substrates. Both products exhibit a very high performance in microwave experiments. After only

5 minutes, 100% of the product is obtained in most experiments. Both also present high selectivity, with yields nearly reaching 100%. The products were also tested for chlorinated substrates as presented on the following page.

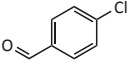
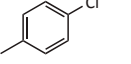
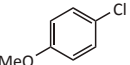
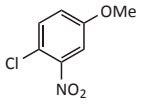
Catalytic Performance in Microwave		
Substrate (R)	Conversion (%) / Yield (%)	
	SiliaCat DPP-Pd (0.5 mol %) <sup>a-b</sup>	SiliaCat S-Pd (0.5 mol %) <sup>c</sup>
	100 / 99.5	100 / 99.3
	100 / 99.4	100 / -
	100 / 88	100 / -
	98 / 97.3	72 / -

General exp. cond.: 1 eq. substrate, 1.1 eq. PhB(OH)<sub>2</sub>, 1.5 eq. K<sub>2</sub>CO<sub>3</sub>; <sup>a</sup> MeOH (0.2 M), 5 min, 75°C, 150 W, 150 psi; <sup>b</sup> MeOH (0.2 M), 5 min, 75°C, 200 W, 200 psi; <sup>c</sup> 15 min.



## Pd-based SiliaCat's Catalytic Performance Comparison (con't)

The SiliaCat Pd-based catalysts can also be used for Suzuki coupling with chlorinated substrates in both conventional and microwave conditions. We have chosen to do this study with SiliaCat DPP-Pd.

Catalytic Performance of Chlorinated Substrates			
Substrate (R)	Conversion /Yield (%) <sup>a</sup>	Substrate (R)	Conversion /Yield (%) <sup>a</sup>
	98 / 93		99 / -
	98 / 96		Bulk: 100 / 98 MW <sup>b</sup> : 100 / 95

<sup>a</sup>Exp. cond. in bulk: 1.5 mol % of SiliaCat, 1 eq. substrate, 1.5 eq. PhB(OH)<sub>2</sub>, 2 eq. K<sub>2</sub>CO<sub>3</sub>, EtOH/H<sub>2</sub>O 15% (0.12M), 6 h, reflux. <sup>b</sup> Microwave: 1 mol % of SiliaCat, 15 min, 125°C.

## Conclusion for Suzuki Coupling

In conclusion, SiliaCat, can be used successfully for Suzuki coupling reactions with iodide, bromide and chloride aryl substrates in conventional or in microwaves conditions. The SiliaCat DPP-Pd gives better performance versus the SiliaCat S-Pd and nearly equivalent to the SiliaCat Pd<sup>0</sup> for the substrates presented.

## Suzuki Coupling Typical Experimental Condition

### Conventional Experimental Conditions

#### Reaction

All products except the catalyst are added to a round bottom flask equipped with a condenser and a magnetic stirrer. Bring mixture to reflux and, after 10 minutes (*when the solution is homogeneous*), add the required quantity of catalyst.

#### Work-up

Once the reaction is complete (*determined by TLC or GC-MS*), the catalyst is filtered at room temperature, rinsed twice with water and three times with the solvent used in the reaction, and finally dried and stored for future use. The reaction mixture obtained after filtration of the catalyst is evaporated, and the product is extracted using ethyl acetate (AcOEt) or diethyl ether (Et<sub>2</sub>O) and washed twice with water. The organic phase is dried using magnesium sulfate (MgSO<sub>4</sub>), and filtered, and the solvent is evaporated. The crude mixture is purified using flash chromatography, if needed. Also applicable to microwave conditions.

### Microwave Conditions

#### Reaction

All products are added to a microwave tube equipped with a magnetic stirrer. Set microwave conditions to:

- Power: 150 W
- Pressure: 150 psi
- Temperature: 75 - 150°C
- Reaction Time: 5 - 15 min

### Suzuki Coupling Typical Experimental Conditions

Products	Conventional Conditions			Microwave Conditions		
	Ar-Iodide	Ar-Bromide	Ar-Chloride	Ar-Iodide	Ar-Bromide	Ar-Chloride
Base [K <sub>2</sub> CO <sub>3</sub> ]	1.5 eq.	1.5 eq.	2.0 eq.	1.5 eq.	1.5 eq.	2.0 eq.
Boronic Acid	1.2 eq.	1.2 eq.	1.5 eq.	1.2 eq.	1.2 eq.	1.5 eq.
SiliaCat Catalyst	≥ 0.5 mol %	≥ 0.5 mol %	≥ 1.0 mol %	≥ 0.5 mol %	≥ 0.5 mol %	≥ 1.0 mol %
Best Solvents (HPLC Grade)	MeOH (0.05 - 0.1 M)	EtOH/H <sub>2</sub> O (10:1, 0.12 M)	EtOH or TBA/H <sub>2</sub> O (10:1.5, 0.12 M)	MeOH (0.2 M)	MeOH (0.2 M)	EtOH/H <sub>2</sub> O (10:1, 0.2 M)

\*Note: molar concentration is related to the substrate.

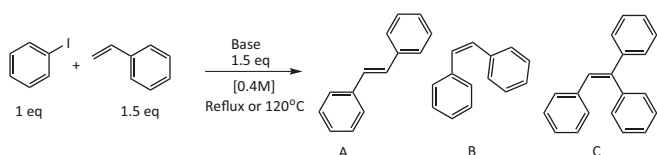
## Heck Coupling Using SiliaCat DPP-Pd & S-Pd

The Heck reaction, also known as the Mizoroki-Heck reaction, is the coupling of a halide with an alkene in the presence of a base and a palladium catalyst. This coupling allows a substitution reaction on alkenes and is of great importance in pharmaceutical research. We determined that the best catalyst for this reaction is SiliaCat DPP-Pd. It showed good reactivity for aryl iodides, bromides and chlorides.

Note: SiliaCat Pd<sup>0</sup> results were not available at the time of printing. Contact us for details.

### Base and Solvent Effects

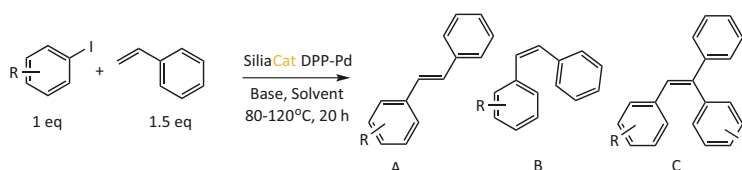
The Heck coupling between iodobenzene and styrene was used to evaluate the influence of solvent and base. The best combinations are KOAc/DMF, Et<sub>3</sub>N/MeCN and Pr<sub>3</sub>N/neat. Using these systems, high yields and great selectivity in favor of product A were obtained.



Base and Solvent Effects (SiliaCat DPP-Pd)				
SiliaCat (mol %)	Base	Solvent (0.4 M)	Time (h)	Conversion A / B / C (%)
0.5	KOAc	DMF	24	100 (90 / 9.5 / 0.5)
0.5	Na <sub>2</sub> CO <sub>3</sub>	DMF	24	67 (62 / 47 / 0)
0.1	Et <sub>3</sub> N	MeCN	24	93 (77 / 6 / 11)
0.1	Et <sub>3</sub> N	H <sub>2</sub> O	24	75 (70 / 5 / 0)
0.1	Pr <sub>3</sub> N	(neat)	20	100 (95 / 5 / 0)

### Catalytic Performance and Comparison vs Homogeneous Catalyst

SiliaCat DPP-Pd is a very efficient catalyst for the Heck coupling and allows greater selectivity over homogeneous Pd catalyst (*TPP is required*). In addition to higher yield of the desired product, the catalyst left minimal residual Pd, TPP or TPPO in solution that would have otherwise required the use of a metal scavenger, chromatography or trituration to remove.



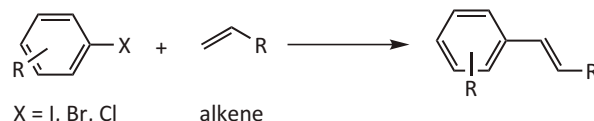
Catalytic Performance and Comparison vs Homogeneous						
Substrate		SiliaCat DPP-Pd (mol %)	Base	Solvent (0.4 M)	Conversion A / B / C (%)	Phosphine Leaching (ppm)
R	X					
4-CN	Br	0.25	NaOAc	DMF	100 (95 / 5 / -)	-
4-NO <sub>2</sub>	Br	0.25	NaOAc	DMF	99 (97 / 2 / -)	-
2-CH <sub>3</sub>	Br	0.25	Et <sub>3</sub> N	MeCN	71 (67 / 5 / -)	-
4-OMe	I	0.25	Et <sub>3</sub> N	MeCN	75 (60 / 15 / -)	-
H	I	0.1	Et <sub>3</sub> N	MeCN	100 (98 / 2 / -)	0
H	I	1.0 Pd(OAc) <sub>2</sub> PPh <sub>3</sub>	Et <sub>3</sub> N	MeCN	100 (70 / 22 / 8)	6,030



## Substrate Scope, Leaching and Microwave Compatibility

SilicaCat catalysts are efficient in the Heck coupling with different substrates. In all cases, conversion and selectivities were excellent. Leaching results were all

below FDA regulations, and no further metal removal was needed. Microwave technology allows faster kinetics with good yields.



Substrate Scope, Leaching and Microwave (MW) Compatibility											
Rn	Mode	SilicaCat DPP-Pd					SilicaCat S-Pd				
		mol %	Time	Temp.	Conv./Sel.(%)	Leaching (ppm)	mol %	Time	Temp.	Conv./Sel.(%)	Leaching (ppm)
1	Batch	0.5	24 h	120°C	100 / 97	-	0.5	24 h	120°C	98 / 92	-
	MW	0.2	10 m	125°C	93 / 85	P: 0.3, Pd: 0.02, Si: 0.8	0.2	15 m	125°C	97 / 93	Pd: 3.8, Si: 1.9
2	Batch	0.2	24 h	135°C	100 / 98	-	0.25	24 h	120°C	85 / 75	-
	MW	0.2 0.5 <sup>1</sup>	10 m 30 m <sup>1</sup>	125°C 150°C <sup>1</sup>	92 / 81 99 / 93 <sup>1</sup>	- P: 0.7, Pd: 0.02, Si: 1.6	0.2	15 m	125°C	87 / 76	Pd: 0.3, Si: 0.8

<sup>1</sup> Et<sub>3</sub>N in water

## Heck Coupling Typical Experimental Procedure

### Conventional Experimental Conditions

#### Reaction

All products except the catalyst are added to a round bottom flask equipped with a condenser and a magnetic stirrer. Bring mixture to reflux (MeCN) or to 120°C (DMF or NMP) and after 10 minutes (*when solution is homogeneous*) add the desired quantity of catalyst.

#### Work-up

Once the reaction is complete (*determined by TLC or GC-MS*), follow the same work-up procedure as for Suzuki coupling conventional experimental conditions as they are applicable to microwave conditions.

### Microwave Conditions

#### Reaction

All products are added to a microwave tube equipped with a magnetic stirrer. Set microwave conditions to:

- Power: 100 W (I-) or 200 W (Br-, Cl-)
- Pressure: 150 psi (I-) or 200 psi (Br-, Cl-)
- Temperature: 100°C (I-) or 125°C (Br-, Cl-)
- Reaction Time: 10 min (I-) or 15 min (Br-, Cl-)

### Heck Coupling Typical Experimental Conditions

Products	Conventional Conditions for 1 eq of:			Microwave Conditions for 1 eq of:		
	Ar-Iodide	Ar-Bromide	Ar-Chloride	Ar-Iodide	Ar-Bromide	Ar-Chloride
Base	1.5 eq. [Et <sub>3</sub> N or NaOAc]	1.5 eq. [NaOAc]	1.5/0.5 eq. [Ca(OH) <sub>2</sub> /TBAB]	1.5 eq. [Et <sub>3</sub> N or NaOAc]	1.5 eq. [K <sub>2</sub> CO <sub>3</sub> ]	2.0 eq. [K <sub>2</sub> CO <sub>3</sub> ]
Olefin	1.2 eq.	1.2 - 2.0 eq.	1.2 - 2.0 eq.	1.2 - 2.0 eq.	1.2 - 2.0 eq.	1.2 - 2.0 eq.
SilicaCat Catalyst	≥ 0.5 mol %	≥ 0.5 mol %	≥ 1.0 mol %	≥ 0.2 mol %	≥ 0.2 mol %	≥ 1.0 mol %
Best Solvents (HPLC Grade)	MeCN (1.2 M) DMF (0.75 M)	DMF (0.75 - 1.5 M)	NMP/H <sub>2</sub> O (1:1, 1.67 M)	MeOH (0.2 M)	MeOH (0.2 M)	EtOH/H <sub>2</sub> O (10:1, 0.2 M)

\*Note: molar concentration is related to the substrate.

## Sonogashira Coupling Using SiliaCat Catalysts

The Sonogashira coupling reaction of aryl halides and terminal acetylenes is an effective method for the formation of substituted acetylenes. This reaction is frequently utilized as a key step in natural product chemistry and for the synthesis of acetylene compounds, which have several applications.

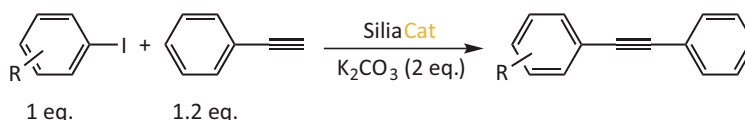
### Catalyst Concentration and Solvent Effects

Sonogashira coupling between iodonitrobenzene and phenylacetylene was achieved easily and without the need for co-catalysts to activate the alkyne, making the use of SiliaCat an efficient method for the formation of substituted acetylenes. All catalysts screened presented excellent efficiency, even in low amounts.

Catalyst Concentration and Solvent Effects														
SilicaCat DPP-Pd					SilicaCat S-Pd					SilicaCat Pd <sup>0</sup>				
mol %	Solvent (M)	Time (min)	Conv. (%)	TON (TOF)	mol %	Solvent (M)	Time (min)	Conv. (%)	TON (TOF)	mol %	Solvent (M)	Time (min)	Conv. (%)	TON (TOF)
0.5	EtOH/H <sub>2</sub> O (0.07)	30	100	200 (400)	0.5	EtOH/H <sub>2</sub> O (0.07)	5	100	200 (2,500)	0.1	EtOH (0.1)	2 h	100	1,000 (500)
0.5	MeOH/H <sub>2</sub> O (0.07)	5	100	200 (2500)	0.5	MeOH/H <sub>2</sub> O (0.07)	1 h	100	200 (200)	0.1	EtOH (0.05)	30	100	1,000 (2,000)
0.1	EtOH/H <sub>2</sub> O (0.07)	1 h	100	1,000 (1,000)	0.1	EtOH/H <sub>2</sub> O (0.07)	1 h	100	1,000 (1,000)					
0.1	MeOH/H <sub>2</sub> O (0.07)	15	100	1,000 (4,000)										
0.01	EtOH/H <sub>2</sub> O (0.13)	3 h	100	10,000 (4,000)										
0.002	EtOH/H <sub>2</sub> O (0.13)	8 h	100	50,000										

### Iodo-Substrate Scope and Microwave Compatibility

Sonogashira couplings between iodoaryls and phenylacetylene are achieved with ease and without the need for co-catalysts to activate the alkyne. This shows that SiliaCat is an efficient tool for the formation of substituted acetylenes.



Iodo- Substrate Scope and Microwave Compatibility										
R	Mode	SilicaCat DPP-Pd			SilicaCat S-Pd			SilicaCat Pd <sup>0</sup>		
		mol %	Conditions	Conv. / Sel. (%)	mol %	Conditions	Conv. / Sel. (%)	mol %	Conditions	Conv. / Sel. (%)
4-NO <sub>2</sub>	Batch	1	EtOH (0.08 M) 77°C, 4 h	100 / 100	1	EtOH (0.08 M) 77°C, 4 h	100 / 100	1	EtOH (0.08 M) 77°C, 4 h	100 / 100
	MW	0.6	MeOH/H <sub>2</sub> O (0.2 M) 100°C, 2 min	100 / -	0.6	MeOH/H <sub>2</sub> O (0.2 M) 100°C, 2 min	100 / -	0.1	MeOH (0.1 M) 75°C, 5 min	100 / -
4-OMe	Batch	1	EtOH (0.08 M) 77°C, 4 h	99 / 98	1	EtOH (0.08 M) 77°C, 4 h	100 / 100	1	EtOH (0.08 M) 77°C, 4 h	99 / 98
4-CH <sub>3</sub>	Batch	1	EtOH (0.08 M) 77°C, 4 h	100 / 100	1	EtOH (0.08 M) 77°C, 4 h	100 / 99	1	EtOH (0.08 M) 77°C, 4 h	100 / 100
	MW	0.5	MeOH/H <sub>2</sub> O (0.2 M) 100°C, 2 min	90 / -	1	MeOH/H <sub>2</sub> O (0.2 M) 100°C, 2 min	88 / -	0.1	MeOH (0.2 M) 75°C, 5 min	100 / -

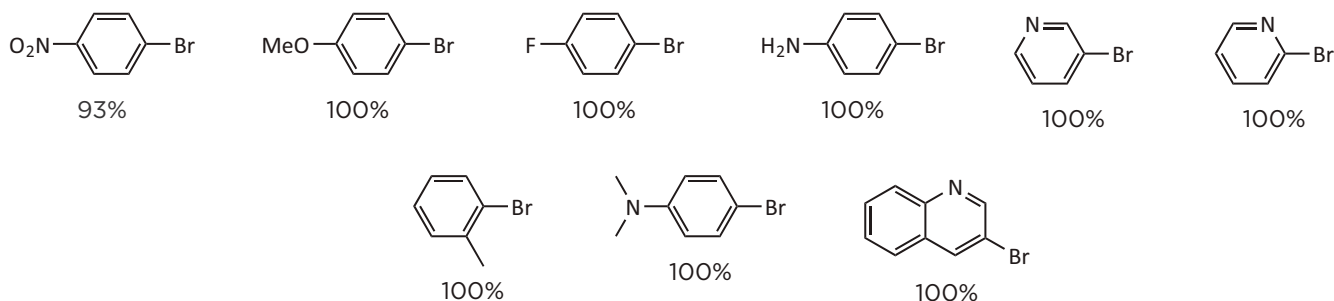


## Bromo- Substrate Scope and Microwave Compatibility

SiliaCat DPP-Pd and Pd<sup>0</sup> are also efficient catalysts for use with bromo substrates. A few examples of the Sonogashira coupling between various bromoaryls substrates (1 eq.) and phenylacetylene (1.25 eq.) using K<sub>2</sub>CO<sub>3</sub> (2 eq.) in MeOH (0.2 M) are shown below.

Conversions obtained with 1 mol % of SiliaCat DPP-Pd under microwave irradiation are presented below. Conventional methodology is also possible, but kinetics are significantly lower (*a few hours compared to 15 minutes*).

### Bromo- Substrate Scope Conversion (%) Results using SiliaCat DPP-Pd



## Sonogashira Coupling Typical Experimental Procedure

### Conventional Experimental Conditions

#### Reaction

All products except the catalyst are added to a round bottom flask equipped with a condenser and a magnetic stirrer. Bring mixture to reflux, and after 10 minutes (*when the solution is homogeneous*) add the required quantity of catalyst.

#### Work-up

Once the reaction is complete (*determined by TLC or GC-MS*), follow the same work-up procedure as for Suzuki coupling conventional experimental conditions as they are applicable to microwave conditions.

### Microwave Conditions

#### Reaction

All products are added to a microwave tube equipped with a magnetic stirrer. Set microwave conditions to:

- Power: 150 W (I-) or 200 W (Br-)
- Pressure: 150 psi (I-) or 200 psi (Br-)
- Temperature: 100°C (I-) or 100 - 150°C (Br-)
- Reaction Time: 5 - 15 min (I-) or 5 - 20 min (Br-)

### Sonogashira Coupling Typical Experimental Conditions

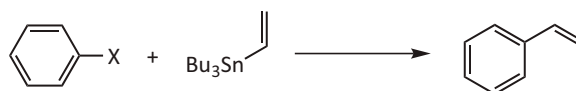
Products	Standard Conditions for 1 eq of:		Microwave Conditions for 1 eq of:	
	Ar-Iodide	Ar-Bromide	Ar-Iodide	Ar-Bromide
Base [K <sub>2</sub> CO <sub>3</sub> ]	1.5 eq.	1.5 eq.	1.5 eq.	2.0 eq.
Alkyne	1.1 eq.	1.25 eq.	1.10 eq.	1.5 eq.
SiliaCat Catalyst	≥ 0.5 mol %	≥ 1.0 mol %	≥ 0.5 mol %	≥ 1.0 mol %
Best Solvents (HPLC Grade)	For room temperature reaction: MeOH (0.02 M) For reflux reaction: MeOH (0.05 - 0.13 M, typ.: 0.07 M) or EtOH/H <sub>2</sub> O (10:1, 0.1 M)		MeOH (0.2M)	MeOH/H <sub>2</sub> O (10:1, 0.2 M)

\*Note: molar concentration is related to the substrate.

## Stille Coupling Using SiliaCat Pd Catalysts

The Stille coupling is a versatile reaction for C-C bond formation. It is a coupling between a halide and an organotin compound. This reaction is widely used in synthesis, but a major drawback is the toxicity of the tin compounds involved. In Stille couplings, a Pd<sup>0</sup> or Pd<sup>II</sup> catalyst is required, and it must be compatible with a wide variety of functional groups (*very few limitation on the R-group*). SiliCycle has developed catalysts that are highly efficient for Stille couplings, as shown below.

Note: SiliaCat Pd<sup>0</sup> results were not available at the time of printing.



### Catalyst Concentration and Solvent Effects

Increasing the amount of the catalyst, for the same solvent and at a constant substrate concentration, improves kinetics (*see table below*). With a mol % of 0.25, the reaction was not completed in 22 h. With a mol % of 2.0, the reaction was completed in 17 h.

As a general rule, if the solvent and the concentration of the substrate are kept constant, increasing the amount of the catalyst, thus increasing the number of the active sites, will speed up the kinetics of the reaction.

This table also shows the importance of the solvent. At low catalyst concentration, 0.25 mol % in dioxane, the reaction was not completed in 22 h. However, under the same conditions but with toluene as the solvent, the reaction was completed in 16 h. In dioxane, the same activity is observed for a

concentration of 2.0 mol %. The solvent is responsible for diffusion of the substrate to the active sites, so the better the diffusion, the higher the kinetics will be.

In all experiments, determining the optimal quantity of SiliaCat in respect to the solvent should be done.

Catalyst Concentration and Solvent Effects

SiliaCat DPP-Pd (mol %)	Solvent (M)	Time (h)	Conversion (%)
2.0	Dioxane (0.1 M)	17	99
0.5	Dioxane (0.1 M)	20	100
0.25	Dioxane (0.1 M)	22	74
0.25	Toluene (0.1 M)	16	99

### SiliaCat DPP-Pd Reusability and Leaching

The minimal leaching and the robustness of the organoceramic matrix are important factors that allow SiliaCat DPP-Pd to be reused several times.



SiliaCat DPP-Pd Reusability and Leaching

Reusability	Conversion (%)	Pd Leaching (ppm)
1 <sup>st</sup>	100	3.0
2 <sup>nd</sup>	100	1.7
3 <sup>rd</sup>	100	2.3
4 <sup>th</sup>	100	2.3





## Catalytic Activity and Additive CsF Influence

Reactions were performed at reflux until the GC/MS analysis showed maximum conversion. Anhydrous conditions are not required.

Catalytic Activity and Additive CsF Influence						
Substrate (R)	Halide (X)	SiliaCat DPP-Pd (mol %)	Additive (eq.)	Solvent (M)	Time (h)	Conversion (%)
4-CN	Br	2	-	Dioxane (0.1 M)	18	87
4-F	Br	10	-	Dioxane (0.1 M)	24	99
4-F	Br	10	CsF (2)	Toluene (0.1 M)	24	100
H	Br	10	-	Toluene (0.1 M)	24	100
4-CH <sub>3</sub>	Br	10	CsF (2)	Dioxane (0.1 M)	24	100
4-OCH <sub>3</sub>	Br	10	CsF (2)	Dioxane (0.1 M)	24	100
H	I	10	CsF (2)	Toluene (0.1 M)	24	100
4-NO <sub>2</sub>	I	2	-	Dioxane (0.1 M)	18	88

Note: R'SnBu<sub>3</sub> was vinyl (1.1 eq.)

## SiliaCat DPP-Pd vs Competitive Catalysts

Comparative analysis with other Pd catalysts available on the market demonstrates the SiliaCat DPP-Pd to be comparable or better in standard Stille conditions. Table at right shows conversion %.

SiliaCat DPP-Pd vs Competitive Catalysts					
SiliaCat DPP-Pd	Escat 1351	Encat 30	Royer Catalyst	Pd(PPh <sub>3</sub> ) <sub>4</sub>	Pd(OAc) <sub>2</sub>
99	44	95	90	72	20

## Stille Coupling Typical Experimental Procedure

### Conventional Experimental Conditions

#### Reaction

All products except the catalyst are added to a round bottom flask equipped with a condenser and a magnetic stirrer. Bring mixture to reflux, and after 10 minutes (*when the solution is homogeneous*) add the required quantity of catalyst.

#### Work-up

Once the reaction is complete (*determined by TLC or GC-MS*), follow the same work-up procedure as for Suzuki coupling standard conditions.

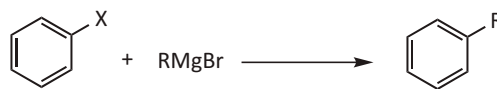
### Experimental Conditions - Stille Coupling

Products	Standard Conditions for 1 eq of: Ar-Iodide & Ar-Bromide
Base [K <sub>2</sub> CO <sub>3</sub> ]	1.0- 2.0 eq. (usually 1.1 eq.)
Additive (CsF)	If needed, add 2.0 eq. for higher conversion
SiliaCat Catalyst	0.25 - 10.0 mol % (typ.: 2 mol % for -I and 2 - 10 mol % for -Br)
Best Solvents (HPLC Grade)	Dioxane (0.1 M) or Toluene (0.1 M)

\*Note: molar concentration is related to the substrate.

## Kumada Coupling Using SiliaCat Pd Catalysts

The Kumada coupling is the direct cross-coupling between an alkyl or an aryl Grignard and a halocarbon. It can be catalyzed by a Pd or a Ni catalyst.

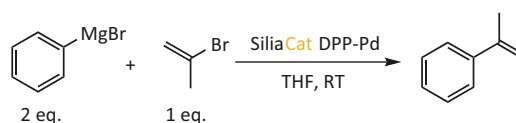
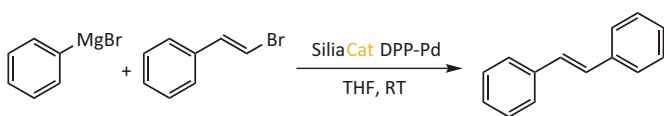


Note: SiliaCat Pd<sup>0</sup> results were not available at the time of printing.

### Catalyst Concentration Effect

At a constant concentration of substrate, an increase of the amount of SiliaCat from 0.1 (or 0.2) to 1.0 will increase the kinetics (*completed in only 15 minutes*).

By increasing the concentration of the catalyst, thus thereby increasing the number of active sites, conversion of the substrate will be favored.



Catalyst Concentration Effect

SiliaCat DPP-Pd (mol %)	Solvent (M)	Time (min)	Conversion (%)
1.0	THF (0.07 M)	15	96
0.5	THF (0.07 M)	15	95
0.2	THF (0.08 M)	2 h	94

Catalyst Concentration Effect

SiliaCat DPP-Pd (mol %)	Solvent (M)	Time (min)	Conversion (%)
1.0	THF (0.08 M)	15	98
0.5	THF (0.08 M)	90	96
0.2	THF (0.08 M)	4 h	98

### Catalyst Reusability and Leaching

Minimal leaching and the robustness of the organoceramic matrix are important factors that allow it to be reused several times.



SiliaCat Reusability and Leaching

Reusability	Conversion (%)	Leaching (ppm)	
		Pd	Si
1 <sup>st</sup>	98	0.20	1.5
2 <sup>nd</sup>	95	0.20	2.3
3 <sup>rd</sup>	94	0.50	1.7
4 <sup>th</sup>	77	0.02	1.9



## Catalytic Activity and Leaching

SiliaCat DPP-Pd showed good reactivity for aryl iodides and bromides. Inert conditions are required for Kumada couplings due to the presence of Grignard reagent.

Catalytic Activity and Leaching						
Substrate (R) / Halide (X)	R-MgBr (2 eq.)	Solvent (M)	Time (h)	Conversion (%)	Leaching (ppm)	
					Pd	Si
4-OCH <sub>3</sub> / Br	Ph-MgBr	THF (0.05 M)	18	98	0.3	0.2
4-OCH <sub>3</sub> / Br	i-Bu-MgBr	THF (0.05 M)	18	95	-	-
4-CH <sub>3</sub> / Br	Ph-MgBr	THF (0.05 M)	18	96	-	-
4-CH <sub>3</sub> / Br	i-Bu-MgBr	THF (0.05 M)	18	98	-	-
4-F / Br	Ph-MgBr	THF (0.08 M)	24	94	< 0.01	1.5
H / I	Ph-MgBr	THF (0.08 M)	24	99	-	-
4-OCH <sub>3</sub> / I	Ph-MgBr	THF (0.08 M)	24	94	-	-
4-CH <sub>3</sub> / I	Ph-MgBr	THF (0.08 M)	24	95	< 0.01	1.0

## Kumada Coupling Typical Experimental Procedure

### Conventional Experimental Conditions

#### Reaction

All products under inert conditions (*catalyst, solvent, substrates, and Grignard reagent*) are added to a Schlenk or a dry round bottom flask equipped with a magnetic stirrer. The mixture was stirred at room temperature until the TLC or GC-MS analysis confirmed reaction completion (18-24h).

#### Work-up

Once the reaction is completed, inert conditions are not necessary. Follow same work-up procedure as for Suzuki coupling conventional experimental conditions.

### Experimental Conditions - Kumada Coupling

Products	Standard Conditions for 1 eq of:
	Ar-Iodide & Ar-Bromide
R-MgBr	2.0 eq.
SiliaCat Catalyst	2.0 - 10.0 mol % (usually 5 mol %)
Best Solvents (HPLC Grade)	Tetrahydrofuran (0.05 - 0.08 M) (usually 0.08 M)

\*Note: molar concentration is related to the substrate. Reaction need to be done at room temperature under inert atmosphere.

## Selective Hydrogenation of Nitroarenes Using SiliaCat Pt<sup>0</sup>

Functionalized anilines are important intermediates in various industries such as pharmaceuticals, polymers, and dyes. Simple aromatic amines are generally obtained by catalytic hydrogenation of nitroarene compounds with various heterogeneous commercial catalysts (*supported nickel, copper, cobalt*) including Pt/C. Yet, the selective reduction of a nitro group with H<sub>2</sub> when other reducible groups are present in the same molecule is

generally not feasible with these catalytic materials and requires the use of advanced heterogeneous catalysts. SiliaCat Pt<sup>0</sup> exhibits chemoselective catalytic activity for the hydrogenation reaction of a series of substituted nitro compounds under remarkably mild conditions, namely at room temperature with 1 bar H<sub>2</sub> in a simple hydrogen balloon, using a modest 0.5 mol % catalyst amount.

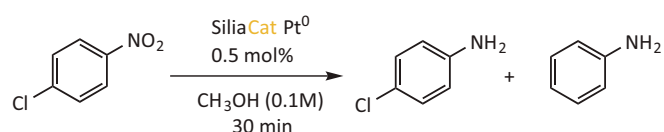
### Solvent and Catalyst Concentration Effects

The best results were obtained using methanol as solvent at 0.1 M concentration with respect to substrate. Even if the use of EtOAc results in high selectivities, the reaction times are generally much longer. Complete conversion is obtained after 1 hour in hexane using 0.5 or 0.1 mol % catalyst, but the selectivity to 4-chloroaniline was generally low.

Solvent and Catalyst Concentration Effects				
SiliaCat Pt <sup>0</sup> (mol %)	Time (h)	Solvent (M)	Yield (%)	
			Product	Aniline
1.0	0.5	MeOH (0.1 M)	92	8
0.5	0.5	MeOH (0.1 M)	87	13
0.2	1	MeOH (0.1 M)	84	13
0.1	2	MeOH (0.1 M)	90	10
0.5	4	EtOAc (0.1 M)	55	0.5
1.0	4	EtOAc (0.1 M)	75	1
1.0	1	THF (0.1 M)	45	17

### SiliaCat Pt<sup>0</sup> Reusability and Leaching

The reusability test of SiliaCat Pt<sup>0</sup> was studied using 4-chloronitrobenzene as substrate under the optimal reaction conditions identified above. Reusing the catalyst in 7 consecutive cycles did not result in any loss of catalytic activity and leaching of Pt and Si (*assessed by ICP-MS*) was minimal. Complete substrate conversion was obtained even after the seventh cycle, with 99% selectivity. The selectivity of the reaction even improves with each subsequent cycle going from 84% in the first run up to 99% in run 7. The positive-feedback phenomenon of effective selectivity in consecutive reaction cycles is probably attributed to the silica matrix alkylation.



SiliaCat Pt <sup>0</sup> Reusability and Leaching				
Reusability	Yield (%)		Leaching (ppm)	
	Product	Aniline	Pt	Si
1	84	12	0.20	1.20
2	89	11	0.04	0.40
3	90	10	0.02	0.08
4	92	8	0.17	0.10
5	98	2	0.01	0.10
6	99	1	0.01	0.12
7	99	1	0.01	0.08



## SiliaCat Pt<sup>0</sup> vs Competitive Catalysts

Other commercially available Pt heterogeneous catalysts [Pt/C, Pt/SiO<sub>2</sub> and Reaxa Pt(O)EnCat40] were tested in the selective reduction of 4-chloro-nitrobenzene. In comparison to other Pt(O) heterogeneous catalysts, the SiliaCat Pt<sup>0</sup>

catalyst proved to be much more reactive, with complete conversion after 0.5 h with just 0.5 mol %. Furthermore, selectivity was significantly higher with only 4% aniline formed as by-product. No secondary product was observed in solution.

SiliaCat Pt<sup>0</sup> vs Competitive Catalysts

Catalyst Mol %	Pt/C			Pt/SiO <sub>2</sub>			Reaxa Pt(O)EnCat40 wet			Reaxa Pt(O)EnCat40 dry		
	5	1	0.5	5	1	0.5	5	1	0.5	5	1	0.5
Time (h)	1	1	1	1	2	2	1	2	2	0.5	2	2
Product (%)	82	65	43	84	88	48	75	78	72	87	90	86
Aniline (%)	14	4	0	13	10	2	18	14	12	13	10	13

Exp. conditions: 2 mol substrate in 20 mL MeOH under hydrogen balloon at room temperature.

## Substrate Scope and Selectivity

The Hydrogenation of different nitro compounds, including those nitro compounds containing different functionalities, was attempted to demonstrate the selectivity of SiliaCat Pt<sup>0</sup> catalyst in a wide range of reactions. The material was tested under hydrogen

balloon, at room temperature conditions in methanol solvent with 0.5 - 1 mol % Pt catalyst.

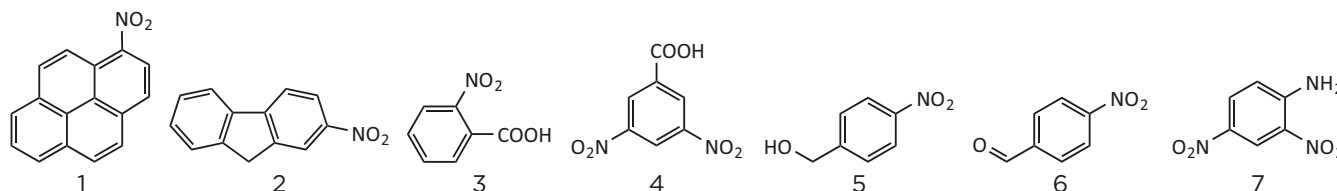
Note: look at our publication in *Adv. Synth. Catal.*, **2011**, 353, 1306-1316 for more examples.

Substrate Scope and Selectivity

Substrate	SiliaCat Pt <sup>0</sup> (mol %)	Solvent (M)	Time (h)	Conversion (%)	Selectivity (%)
Structure #1	0.5	MeOH (0.05 M)	1	100	98 (5% pyrene)
Structure #2	0.5	MeOH (0.05 M)	1	98	100
Structure #3	0.5	MeOH (0.1 M)	1	100	100
Structure #4	0.5 / 1.0	MeOH (0.1 M)	2	100 / 100 <sup>1</sup>	98 / 100 <sup>1</sup>
Structure #5	0.5	MeOH (0.1 M)	1	100	100
Structure #6	0.5	MeOH (0.1 M)	1	100	95
Structure #7	0.5	MeOH (0.07 M)	2	100	100

<sup>1</sup> If 0.5 mol % was used only one NH<sub>2</sub> group was reduced. If 1 mol % was used, both nitro groups were reduced.

## Substrate Structures



## Conclusion of Selective Hydrogenation of Nitroarenes

The hydrogenation of different nitro compounds and the selective hydrogenation of different nitro compounds in the presence of different functionalities, including reducible carbonyl, amide, ester, amine and halide groups was achieved with SiliaCat Pt<sup>0</sup> catalyst in methanol at room temperature and under 1 bar H<sub>2</sub> pressure. Given the broad applicability of Pt-based catalysts to widely different chemical reactions, it is envisaged that these catalysts, now commercially available, will be used in numerous fields of chemical synthesis as well as in energy generation applications.

## Selective Hydrogenation of Nitroarenes Typical Experimental Procedure

### Conventional Experimental Conditions

#### Reaction

Typical reactions are performed on a 2 mmol scale. The substrate is dissolved in 20 mL of MeOH and then treated with 0.1 - 1 mol % of SiliaCat Pt<sup>0</sup> catalyst. The mixture is degassed twice, replacing the vacuum by hydrogen each time. The reaction mixture, connected to a balloon of hydrogen, is stirred at room temperature until it shows maximum conversion.

#### Work-up

Once the reaction is complete (*determined by TLC or GC-MS*), the catalyst is filtered off and washed with EtOH or MeOH. The filtrate is concentrated to give a crude product, and the conversion to the desired product is determined by GC/MS analysis.

#### Reusability

To reuse the catalyst, after completion of the reaction remove the catalyst by filtration, rinse with MeOH/THF solvents and dry under vacuum.

## Selective Debenzylation Using SiliaCat Pd<sup>0</sup>

The selective debenzylation of aryl benzyl ethers, benzyl esters, and benzyl amines, while leaving other sensitive groups intact, can be carried out in high yield under remarkably mild conditions (*namely at room temperature under 1 bar H<sub>2</sub> in a simple hydrogen balloon, using a modest 0.5 mol % catalyst amount*) using SiliaCat Pd<sup>0</sup> (*note that SiliaCat Pt<sup>0</sup> can also be used but reaction times are longer and concentrations are higher*).

Selective and smooth deprotection is critical. The commonly used method makes use of catalytic hydrogenolysis to protect benzylic

groups with H<sub>2</sub> under pressure and in the presence of a heterogeneous catalyst such as Pd/C or Raney Ni. Often, however, the deprotection reaction conditions are not compatible with other functional groups, such as nitro, unsaturated bonds, and halogen groups.

SiliaCat hydrogenation catalysts offer a number of additional advantages over traditional Pd/C. They are non pyrophoric, and have a higher density and lower catalytic consumption (<1 mol % vs 5 - 10% for Pd/C) due to fast kinetics and high turnover.

Note: refer you to our publication called "Selective Debenzylation of Benzyl Protected Groups with SiliaCat Pd<sup>0</sup> under Mild Conditions" in *ChemCatChem*, 2011, 3, 1-5.



## Solvent Effect

Solvent choice is critical for any debenzoylation reaction. Therefore, in order to optimize the reaction conditions, 1-(benzyloxy)-4-methoxybenzene was used as our substrate of choice. A series of commonly employed solvents (*THF, methanol, ethanol, ethyl acetate, and hexane*) were screened under a hydrogen balloon at room temperature and at different solvent concentrations. The best results were achieved with methanol and ethanol (*HPLC grade*).



Solvent Effect			
SiliaCat Pd <sup>0</sup> (mol %)	Time (h)	Solvent (M)	Conversion (%)
2	16	EtOH (0.1 M)	17
2	16	MeOH (0.1 M)	15
2	4	EtOH (0.07 M)	100
2	0.5	MeOH (0.07 M)	100
2	20	THF (0.07 M)	10
2	20	THF (0.07 M)	15
2	20	EtOAc (0.07 M)	20
2	20	Hexane (0.07 M)	21

## Catalyst Concentration Effect

The molar concentration of the solvent with respect to the substrate is crucial with higher concentrations, slowing or even preventing reaction. The best results were achieved by using a methanol concentration of 0.07 M and 0.5 – 1 mol % SiliaCat Pd<sup>0</sup>, with complete conversion obtained after 1 – 2 hours.

Catalyst Concentration Effect			
SiliaCat Pd <sup>0</sup> (mol %)	Time (h)	Solvent (M)	Conversion (%)
2	0.5	MeOH (0.07 M)	100
1	1	MeOH (0.07 M)	100
0.5	2	MeOH (0.07 M)	100

## SiliaCat Pd<sup>0</sup> Reusability and Leaching

Catalyst stability and reusability are crucial features of any catalyst seeking commercial applications. The SiliaCat Pd<sup>0</sup> was thus reused six consecutive times in the O-debenzylation reaction of 1-(benzyloxy)-4-methoxybenzene under the standard mild conditions developed in our laboratory (*reaction shown on previous page using 1 mol % of catalyst*).

After six runs, the catalyst exhibits only a slight loss in activity compared with that of a catalyst run three times. However, the activity remained approximately constant, and it was enough to expand the reaction time to 1 h and 30 min to gain complete debenzoylation of the substrate.

SiliaCat Pd <sup>0</sup> Reusability and Leaching				
Reusability	Time (h)	Conversion (%)	Leaching (ppm)	
			Pd	Si
1 <sup>st</sup>	1	100	0.7	2.5
2 <sup>nd</sup>	1	100	0.3	1.3
3 <sup>rd</sup>	1	95	-	-
	1.5	100	0.3	2.3
4 <sup>th</sup>	1	95	-	-
	1.5	100	0.2	1.4
5 <sup>th</sup>	1	94	-	-
	1.5	99	0.2	0.8
6 <sup>th</sup>	1	94	-	-
	1.5	100	0.1	0.5

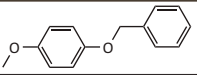
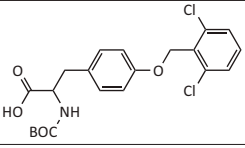
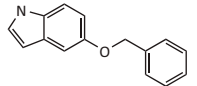
## SiliaCat Pd<sup>0</sup> vs a Competitive Catalyst

Using the same reaction as that used to demonstrate the reusability of SiliaCat Pd<sup>0</sup> in the O-debenzylation reaction of 1-(benzyloxy)-4-methoxybenzene, we also tested the commercial catalyst Pd<sup>0</sup> EnCat, a polyurea-entrapped catalyst.

SiliaCat Pd <sup>0</sup> vs a Competitive Catalyst			
Catalyst (mol %)	Time (h)	Conversion (%)	Selectivity (%)
SiliaCat Pd <sup>0</sup> (0.5)	1 / 2	95 / 100	- / 100
SiliaCat Pd <sup>0</sup> (1.0)	0.5 / 1	75 / 100	- / 100
Pd <sup>0</sup> EnCat (10)	16	100	100

## Substrate Scope and Selectivity

SiliaCat Pd<sup>0</sup> is an efficient catalyst for the selective debenzylation of different aryl benzyl ethers, benzyl amino-acids, and benzylprotected sugars leaving other sensitive groups intact. Refer to our publication in *ChemCatChem*, 2011, 3, 1-5 for more examples.

Substrate Scope & Selectivity					
Substrate	SiliaCat Pd <sup>0</sup> (mol %)	Time (h)	Conversion [Yield] (%)	Leaching (ppm)	
				Pd	Si
	1	1	100 [99.7]	3.4	1.3
	1	1	100 [98.6]	1.7	5.0
	1	20	100 [98.0]	0.4	7.0

## Conclusion of Selective Debenzylation

In conclusion, the SiliaCat Pd<sup>0</sup> catalyst is suitable for the selective debenzylation of numerous substrates under mild conditions with a modest 1 - 2 mol % catalyst amount. Benzyl-protected sugars, amino acids, ethers, and esters are smoothly debenzylated under 0.1 MPa H<sub>2</sub> at room temperature.

## Selective Debenzylation Typical Experimental Procedure

### Conventional Experimental Conditions

#### Reaction

Typical reactions were performed on a 1 mmol scale. The substrate was dissolved in 15 mL of MeOH or EtOH (0.07 M) and 1 or 2 mol % of the SiliaCat Pd<sup>0</sup> catalyst was added. The mixture was degassed twice and each time replacing the vacuum by hydrogen. The reaction mixture, connected to a balloon filled with hydrogen, was stirred at room temperature until GC/MS analysis showed maximum conversion.

#### Reusability

To reuse the catalyst, after completion of the reaction, remove the catalyst by filtration, rinse with MeOH/THF solvents and dry under vacuum.

#### Work-up

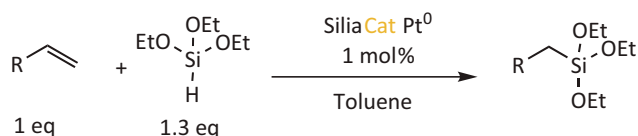
Once the reaction is complete as deemed by TLC or GC-MS, the catalyst was filtered off and washed with EtOH or MeOH. The filtrate was concentrated to give a crude product. The conversion in the desired product was determined by GC/MS analysis and by <sup>1</sup>H NMR.





## Hydrosilylation Using SiliaCat Pt<sup>0</sup>

Hydrosilylation reactions (*or catalytic hydrosilylation*) are a widely used method to prepare organosilicon products. The reaction consists of the addition of Si-H bonds on unsaturated bonds like alkenes, alkynes or ketones, where catalysts are often required (usually H<sub>2</sub>PtCl<sub>6</sub>). SiliaCat Pt<sup>0</sup> can be used for hydrosilylation reactions. Some examples are shown to the right.



Hydrosilylation using SiliaCat Pt<sup>0</sup>

Substrate	Time (h)	Temp. (°C)	Conversion (%)	Selectivity (%)
1-octene	5	22 / 60	88 / 99	98 / 99
1-decene	5	22 / 60	100 / 100	97 / 98
1-octadecene	5	22 / 60	95 / 98	56 / 83
4-vinyl-benzamine	5 / 24	60	47 / 80	96 / 96
3,3-diethoxy-prop-1-ene	5	22 / 60	94 / 100	93 / 81

## Hydrosilylation Typical Experimental Procedure

### Conventional Experimental Conditions

#### Reaction

A 100 ml two neck dry round bottom flask equipped with a condenser and a rubber septum is filled with 1 mol % SiliaCat Pt<sup>0</sup> and was degassed two times for 15 minutes kept under argon conditions. The anhydrous solvent, the silane (95% pure) and the olefin (previously degassed for 15 minutes under argon) were added using a syringe. The reaction mixture was either stirred at room temperature or heated at 60°C until the GC/MS analysis showed maximum conversion.

Note: Unless otherwise indicated, all manipulations were carried out under argon conditions. In general, reactions were performed on a 2 mmol scale in 15 ml anhydrous toluene.

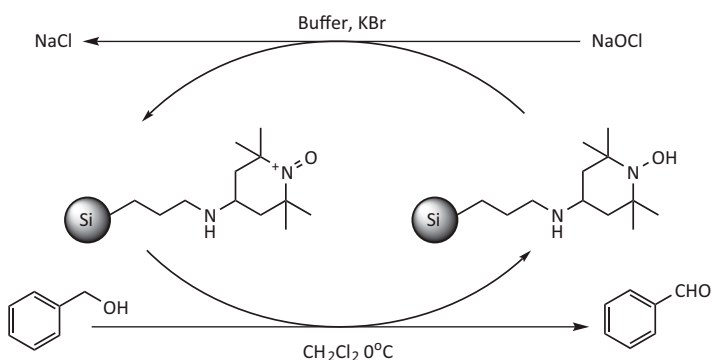
#### Work-up

Once the reaction is complete the catalyst was filtered off and washed with toluene. The filtrate was concentrated to give a crude product. The conversion in the desired product was determined by GC/MS.

## Oxidation Using SiliaCat TEMPO

Aldehydes and ketones, either as starting materials, synthetic intermediates, or final products, are of great interest in synthetic chemistry. Such carbonyl-containing products can lead to carbon-carbon (*i.e.* Wittig, Aldol, alkylation) or carbon-nitrogen bond formation. Over the years, chemists have discovered various oxidizing agents such as pyridinium chlorochromate (PCC), MnO<sub>2</sub>, Dess-Martin periodinane, or Swern oxidation conditions. Although all these methods lead to the aldehyde (*limited oxidation of the aldehyde to the carboxylic acid*), they have drawbacks such as the hazards and toxicity associated with residual metal contamination.

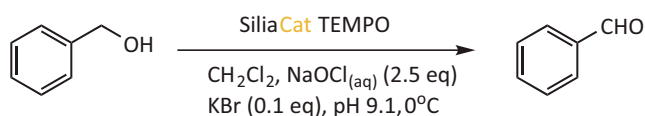
Development of environmentally friendly methods such as selective catalytic oxidation of alcohol substrates to aldehydes and ketones can have significant impact on modern methods of chemical synthesis. SiliaCat TEMPO is the oxidation solution of choice.



### Catalytic Performance and Leaching

SiliaCat TEMPO was investigated in the Montanari-Anelli conditions. The catalytic cycle involves regeneration of the oxidative species with NaOCl (*commercially available bleach*) in presence of KBr as co-catalyst to form the stronger anion OBr<sup>-</sup>.

Unless otherwise stated, the reaction shown below was used for the demonstration.



Catalytic Performance and Leaching				
SiliaCat (mol %)	Time (h)	Conversion (%)	TON	Si Leaching (ppm)
0.1	1	95	950	-
0.01	2	83	8,300	3
0.01	3	95	9,500	1.6
0.01	4	97	9,650	1.5
0.02	2	96	4,800	-
0.02	3	100	5,000	2

SiliaCat TEMPO can be used with as low as 0.01 mol % quantity to provide the desired aldehyde in short reaction times. ICP analysis confirms that the material is leach-resistant ([Si] ≥ 3 ppm).

### SiliaCat TEMPO Reusability

Minimal leaching and the robustness of SiliaCat TEMPO's organoceramic matrix allow it to be reused several times for further uses.

SiliaCat TEMPO Reusability								
Reusability	Time (min)	Conversion (%)	Reusability	Time (min)	Conversion (%)	Reusability	Time (min)	Conversion (%)
1 <sup>st</sup>	30	100	...	...	...	9 <sup>th</sup>	30 / 60	97 / 100
2 <sup>nd</sup>	30	100	8 <sup>th</sup>	30 / 60	95 / 100	10 <sup>th</sup>	30 / 60	90 / 100

<sup>a</sup>SiliaCat TEMPO is recycled by post-reaction filtration, DCM washes and air drying.



## Influence of Co-Catalyst KBr and Temperature

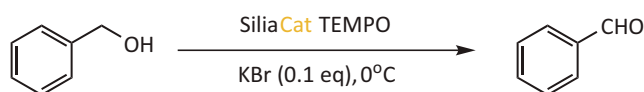
SiliaCycle investigated whether it was necessary to use a co-catalyst (KBr) for the reaction to proceed effectively. As shown in the table, although KBr is not required for the reaction, it does have a significant impact on the kinetics. The reaction can still proceed to completion without KBr but requires longer time and/or more SiliaCat TEMPO. It was also demonstrated that the reaction can be carried out at room temperature without KBr.

**Influence of KBr and Temperature**

SiliaCat (mol %)	KBr (eq.)	Temp. (°C)	Time (min)	Conversion (%)
0.1	0.1	0	60	95
0.1	0	0	60	80
0.1	0	0	210	100
0.2	0	0	105	96
0.2	0	22	60	76
0.2	0	22	90	87

## Influence of Solvents, pH and NaOCl

As shown on the right, the reaction can be carried out at pH 9.0 or at pH 7.5 in DCM with high conversion yields. The catalytic conditions are selective towards the aldehyde, rather than the carboxylic acid, even with 10 equiv of NaOCl. At pH 7 in water, the reaction is slower, but this can be overcome by using more NaOCl<sub>(aq)</sub>. At pH 9, the conversion is high, but too much bleach and the long reaction time in the aqueous media will lead to the corresponding carboxylic acid. The reaction can also be pursued in other organic solvents.



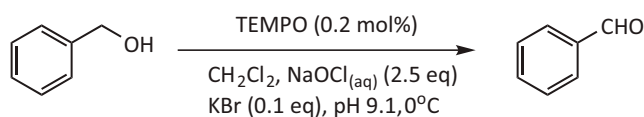
**Influence of Solvent, pH and NaOCl<sub>(aq)</sub>**

SiliaCat (mol %)	NaOCl <sub>(aq)</sub> (eq.)	Solvent	pH	Time (min)	Conversion (%)
0.2	2.50	DCM	9.0	60	98
0.2	10.00	DCM	9.0	90	98
0.2	1.25	DCM	7.5	60 / 90	83 / 86
0.2	2.50	DCM	7.5	60 / 90	94 / 98
0.2	1.25	H <sub>2</sub> O	7.5	60 / 90	57 / 65
0.2	2.50	H <sub>2</sub> O	7.5	60 / 90	87 / 88
0.7	1.20	H <sub>2</sub> O	9.0	60 / 150	83 / 89
0.8	5.00	H <sub>2</sub> O	9.0	60 / 18 h	60 (19) / 7 (89) <sup>1</sup>
0.2	1.25	EtOAc	9.0	60 / 90	95 / 96

<sup>1</sup> In parenthesis = conversion to carboxylic acid.

## SiliaCat TEMPO vs Homogeneous TEMPOs

Comparative analysis versus homogeneous TEMPOs demonstrates the SiliaCat TEMPO to be comparable or better at neutral pH and significantly superior in basic conditions.



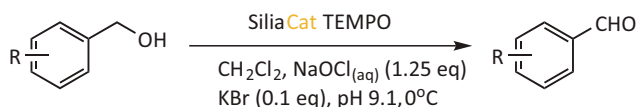
**SiliaCat TEMPO vs Homogeneous TEMPOs**

pH	SiliaCat TEMPO	4-MeO-TEMPO	4-Oxo-TEMPO
7.5	91	99	45
9.0	98	55 (40) <sup>1</sup>	73

<sup>1</sup> In parenthesis = conversion to carboxylic acid.

## Substrate Scope with SiliaCat TEMPO

SiliaCat TEMPO is efficient with different substrates and can be used with phase transfer agents such as Aliquat 336. When an electron-rich benzylic alcohol cannot be oxidized successfully with NaOCl, conditions involving I<sub>2</sub> in toluene, at room temperature, will yield the desired product.



Substrate Scope with SiliaCat TEMPO

Substrate (R)	Catalyst (mol %)	Time (min)	Conversion (%)
3-NO <sub>2</sub>	0.4	90	100
4-NO <sub>2</sub>	0.4	90	98
4-OCH <sub>3</sub>	0.4	90	36
4-OCH <sub>3</sub>	0.4 (0.05 eq. Aliquat 336)	60	79
4-Cl	0.4	90	95
3-phenyl-1-propanol	0.4	60	97
1-phenyl-3-propanol	0.4	180	95
4-OCH <sub>3</sub>	8.2	16 h	99 <sup>1</sup>
3-OCH <sub>3</sub>	7.8	16 h	96 <sup>1</sup>
Piperonal	10.0	20 h	100 <sup>1</sup>

<sup>1</sup>Exp. Cond.: I<sub>2</sub> (1.8 eq.), NaHCO<sub>3</sub>(aq), pH 8, toluene, 22°C.

## Conclusion of Oxidation

In conclusion, the SiliaCat TEMPO is an effective oxidizing catalyst presenting unique advantages such as high activity, robustness, leach-proof properties and selectivity toward the oxidation of alcohols into aldehydes and ketones, both very valuable products in organic chemistry.

## Oxidation Typical Experimental Procedure

### Oxidation of Alcohols or Aldehydes to Carboxylic Acid

Note: changing the solvent to water, increasing temperature and the amount of bleach will all favor the acid formation.

#### Conventional Experimental Conditions

##### Reaction

Under mechanical agitation, a 0.4M solution of alcohol in water and a 0.5 M aqueous solution of KBr were cooled at 0°C in an ice bath. The desired amount of SiliaCat TEMPO was added, followed by an aqueous solution of NaOCl (from 10-13% bleach) buffered at pH 9 (using NaHCO<sub>3</sub>) or pH 6.7 (using NaH<sub>2</sub>PO<sub>4</sub>/Na<sub>2</sub>HPO<sub>4</sub>). NaOCl was added slowly over a 10 minute period as the reaction is exothermic. The mixture was warmed to room temperature (20°C) and stirred between 1,300-1,500 rpm. The temperature can be increased to 35°C if necessary.

- 1.2 - 5 eq. of NaOCl<sub>(aq)</sub> (typically start with 3 eq. and, if necessary, add another 2 eq. of NaOCl via an addition funnel after all of alcohol is consumed)
- 0.1 eq. of potassium bromide (KBr) (prepared as a 0.5 M solution)
- pH 9 is achieved using a NaHCO<sub>3</sub> buffer or a pH of 6.7 is achieved using a sodium phosphate buffer (1:1 mixture of 0.67 M NaH<sub>2</sub>PO<sub>4</sub> and 0.67 M Na<sub>2</sub>HPO<sub>4</sub>)
- 0.01 - 1 mol % of SiliaCat TEMPO (typically 1 mol %)
- The best solvents are H<sub>2</sub>O, ACN/H<sub>2</sub>O or DCM/H<sub>2</sub>O, typically at 0.4 M (molar concentration with respect to the substrate)

##### Work-up

Once the reaction was complete (determined by TLC or GC-MS), the catalyst was filtered at room temperature, and the pH was adjusted to 12 with aqueous NaOH (2N). The aqueous phase was separated, acidified with HCl 6N and extracted with CH<sub>2</sub>Cl<sub>2</sub>. The organic phase was dried over MgSO<sub>4</sub> and evaporated. The residue was purified by crystallization or column chromatography on silica gel.



## Oxidation of Primary or Secondary Alcohols

### Under Montanari-Anelli Conditions (using NaOCl)

#### Conventional Experimental Conditions

##### Reaction

Under mechanical agitation, a 0.4M solution of the alcohol in dichloromethane is mixed with a 0.5M aqueous solution of KBr and cooled at 0°C in an ice bath. The desired amount of SiliaCat TEMPO is then added, followed by an aqueous solution of NaOCl (*from commercially available 10-13% bleach*), then the solution is buffered at pH 9 (*using NaHCO<sub>3</sub>*). NaOCl solution is added slowly over a 10 minute period as the reaction is exothermic. The mixture is then stirred between 1,300-1,500 rpm.

##### Work-up

Once the reaction is complete (*determined by TLC or GC-MS*), the catalyst is filtered at room temperature, and the organic phase is dried over MgSO<sub>4</sub> and evaporated. Crude mixture is purified using flash chromatography, if needed.

### Under Miller Conditions (using I<sub>2</sub> co-catalyst)

#### Conventional Experimental Conditions

##### Reaction

Under mechanical agitation, a 0.4M solution of alcohol in toluene is mixed at room temperature (20°C) with a 0.3 M aqueous solution of NaHCO<sub>3</sub>. Solid iodine is then added in one portion to the mixture, followed by the desired amount of SiliaCat TEMPO.

##### Work-up

Once the reaction is complete (*determined by TLC or GC-MS*), the catalyst is filtered at room temperature. The mixture should then be cooled to 5°C, diluted with ethyl acetate, and quenched with a 0.8M aqueous solution of Na<sub>2</sub>SO<sub>3</sub>. The uncolored organic phase is then washed with a saturated aqueous solution of NaHCO<sub>3</sub> followed by brine and dried over MgSO<sub>4</sub>. After filtration and evaporation of the solvents, the crude mixture can be purified using flash chromatography.

- For Montanari-Anelli conditions: 1.2 - 5 eq. of NaOCl<sub>(aq)</sub> (*typically 2.5 eq.*) and 0.1 eq. of KBr (*prepared as a 0.5 M solution*)
- For Miller conditions: 1.8 eq. of solid iodine (I<sub>2</sub>)
- 0.001 - 1 mol % of SiliaCat TEMPO (*typically 1 mol %*)
- The best solvents are DCM, EtOAc or ACN/H<sub>2</sub>O (*HPLC grade*), typically at 0.4 M (*molar concentration is with respect to the substrate*)

### SiliCycle Publications

#### SiliaCat TEMPO Oxydation

- Topics in Catalysis*, **2010**, 53, 1110-1113
- Organic Process Research & Development*, **2010**, 14, 245-251
- Chemistry Today*, **2009**, 27, 13-16
- Organic Process Research and Development*, **2007**, 11, 766-768

#### Hydrogenation of nitroarenes with SiliaCat Pt<sup>0</sup>

- Advanced Synthesis & Catalysis*, **2011**, 353, 1306-1316
- Catal. Sci. Technol.*, **2011**, Advance Article, DOI: 10.1039/C1CY00097G

#### Suzuki coupling with SiliaCat

- Catal. Sci. Technol.*, **2011**, Advance Article, DOI: 10.1039/C1CY00119A
- Topics in Catalysis*, **2010**, 53, 1059-1062

#### Selective debenzoylation with SiliaCat Pd<sup>0</sup>

- ChemCatChem*, **2011**, 3, 1146-1150